

Note from Mr. Groves... CHECK your work... don't just COPY the answers.

	⑧ K_2CrO_4	⑦ $CuSO_4$	⑥ $Ba(OH)_2$	⑤ $Pb(NO_3)_2$	④ Na_2CO_3	③ KCl	② $(NH_4)_2SO_4$
① $AgNO_3$	Ag_2CrO_4 KNO_3	$Cu(NO_3)_2$ Ag_2SO_4	$Ba(NO_3)_2$ $AgOH$	$AgNO_3$ $Pb(NO_3)_2$	Ag_2CO_3 $NaNO_3$	KNO_3 $AgCl$	Ag_2SO_4 NH_4NO_3
② $(NH_4)_2SO_4$	$(NH_4)_2CrO_4$ K_2SO_4	$CuSO_4$ $(NH_4)_2SO_4$	$BaSO_4$ NH_4OH	NH_4NO_3 $PbSO_4$	Na_2SO_4 $(NH_4)_2CO_3$	K_2SO_4 NH_4Cl	
③ KCl	K_2CrO_4 KCl	$CuCl_2$ K_2SO_4	$BaCl_2$ KOH	KNO_3 $PbCl_2$	$NaCl$ K_2CO_3		
④ Na_2CO_3	Na_2CrO_4 K_2CO_3	$CuCO_3$ Na_2SO_4	$BaCO_3$ $NaOH$	$PbCO_3$ $NaNO_3$			
⑤ $Pb(NO_3)_2$	$PbCrO_4$ KNO_3	$Cu(NO_3)_2$ $PbSO_4$	$Ba(NO_3)_2$ $Pb(OH)_2$				
⑥ $Ba(OH)_2$	$BaCrO_4$ KOH	$Cu(OH)_2$ $BaSO_4$					
⑦ $CuSO_4$	$CuCrO_4$ K_2SO_4						

Follow Up Answers

1 & 8

- $2 \text{AgNO}_3(\text{aq}) + \text{K}_2\text{CrO}_4(\text{aq}) \rightarrow \text{Ag}_2\text{CrO}_4(\text{s}) + 2 \text{KNO}_3(\text{aq})$
- $2 \text{Ag}^+ + 2 \text{NO}_3^- + 2 \text{K}^+ + \text{CrO}_4^{2-} \rightarrow \text{Ag}_2\text{CrO}_4(\text{s}) + 2 \text{K}^+ + 2 \text{NO}_3^-$
- $2 \text{Ag}^+ + \text{CrO}_4^{2-} \rightarrow \text{Ag}_2\text{CrO}_4(\text{s})$

1 & 6

- $2\text{AgNO}_3(\text{aq}) + \text{Ba}(\text{OH})_2(\text{aq}) \rightarrow 2\text{AgOH}(\text{s}) + \text{Ba}(\text{NO}_3)_2(\text{aq})$
- $2\text{Ag}^+ + 2\text{NO}_3^- + \text{Ba}^{2+} + 2\text{OH}^- \rightarrow 2\text{AgOH}(\text{s}) + \text{Ba}^{2+} + 2\text{NO}_3^-$
- $\text{Ag}^+ + \text{OH}^- \rightarrow \text{AgOH}(\text{s})$

1 & 4

- $2\text{AgNO}_3(\text{aq}) + \text{Na}_2\text{CO}_3(\text{aq}) \rightarrow \text{Ag}_2\text{CO}_3(\text{s}) + 2\text{NaNO}_3(\text{aq})$
- $2\text{Ag}^+ + 2\text{NO}_3^- + 2\text{Na}^+ + \text{CO}_3^{2-} \rightarrow \text{Ag}_2\text{CO}_3(\text{s}) + 2\text{Na}^+ + 2\text{NO}_3^-$
- $2\text{Ag}^+ + \text{CO}_3^{2-} \rightarrow \text{Ag}_2\text{CO}_3(\text{s})$

1 & 3

- $\text{AgNO}_3(\text{aq}) + \text{KCl}(\text{aq}) \rightarrow \text{AgCl}(\text{s}) + \text{KNO}_3(\text{aq})$
- $\text{Ag}^+ + \text{NO}_3^- + \text{K}^+ + \text{Cl}^- \rightarrow \text{AgCl}(\text{s}) + \text{K}^+ + \text{NO}_3^-$
- $\text{Ag}^+ + \text{Cl}^- \rightarrow \text{AgCl}(\text{s})$