

## 9 • Bonding & Molecular Structure

### LEWIS STRUCTURES

Indicate the # of **VALENCE** electrons for each species. Write the correct Lewis electron-dot structure for each.

<b>F</b> # of valence e <sup>-</sup> 's = ____	<b>O</b> # of valence e <sup>-</sup> 's = ____	<b>K</b> # of valence e <sup>-</sup> 's = ____	<b>Al</b> # of valence e <sup>-</sup> 's = ____
<b>F<sup>-</sup></b> # of valence e <sup>-</sup> 's = ____	<b>O<sup>2-</sup></b> # of valence e <sup>-</sup> 's = ____	<b>K<sup>+</sup></b> # of valence e <sup>-</sup> 's = ____	<b>Al<sup>3+</sup></b> # of valence e <sup>-</sup> 's = ____
<b>F<sub>2</sub></b> # of valence e <sup>-</sup> 's = ____	<b>H<sub>2</sub></b> # of valence e <sup>-</sup> 's = ____	<b>HF</b> # of valence e <sup>-</sup> 's = ____	<b>NH<sub>3</sub></b> # of valence e <sup>-</sup> 's = ____
<b>CH<sub>4</sub></b> # of valence e <sup>-</sup> 's = ____	<b>NF<sub>3</sub></b> # of valence e <sup>-</sup> 's = ____	<b>SiF<sub>4</sub></b> # of valence e <sup>-</sup> 's = ____	<b>C<sub>2</sub>H<sub>6</sub></b> # of valence e <sup>-</sup> 's = ____
<b>MgH<sub>2</sub></b> # of valence e <sup>-</sup> 's = ____	<b>LiH</b> # of valence e <sup>-</sup> 's = ____	<b>AlH<sub>3</sub></b> # of valence e <sup>-</sup> 's = ____	<b>BH<sub>3</sub></b> # of valence e <sup>-</sup> 's = ____

<b>C<sub>2</sub>H<sub>4</sub></b> # of valence e <sup>-</sup> 's = ____	<b>C<sub>2</sub>F<sub>4</sub></b> # of valence e <sup>-</sup> 's = ____	<b>CO</b> # of valence e <sup>-</sup> 's = ____	<b>O<sub>2</sub></b> # of valence e <sup>-</sup> 's = ____
<b>CO<sub>2</sub></b> # of valence e <sup>-</sup> 's = ____	<b>C<sub>2</sub>H<sub>2</sub> (H C C H)</b> # of valence e <sup>-</sup> 's = ____	<b>N<sub>2</sub></b> # of valence e <sup>-</sup> 's = ____	<b>HCN</b> # of valence e <sup>-</sup> 's = ____
<b>CN<sup>-</sup></b> # of valence e <sup>-</sup> 's = ____	<b>SO<sub>4</sub><sup>2-</sup></b> # of valence e <sup>-</sup> 's = ____	<b>PO<sub>4</sub><sup>3-</sup></b> # of valence e <sup>-</sup> 's = ____	<b>ClO<sub>3</sub><sup>-</sup></b> # of valence e <sup>-</sup> 's = ____
<b>CO<sub>3</sub><sup>2-</sup></b> # of valence e <sup>-</sup> 's = ____	<b>NO<sub>3</sub><sup>-</sup></b> # of valence e <sup>-</sup> 's = ____	<b>SO<sub>2</sub></b> # of valence e <sup>-</sup> 's = ____	<b>O<sub>3</sub> (O O O)</b> # of valence e <sup>-</sup> 's = ____
<b>SF<sub>6</sub></b> # of valence e <sup>-</sup> 's = ____	<b>XeF<sub>4</sub></b> # of valence e <sup>-</sup> 's = ____	<b>PCl<sub>5</sub></b> # of valence e <sup>-</sup> 's = ____	<b>SeF<sub>4</sub></b> # of valence e <sup>-</sup> 's = ____