South Pasadena • AP Chemistry

Name _____
Period ___ Date ___/___/___

21 • Electron Transfer Reaction

PRACTICE TEST

1. Which of the following is the correct cell notation for the reaction

$$Hg_2^{2+} + Cd(s) \rightarrow Cd^{2+} + 2Hg(l)$$

a) $Cd^{2+} \mid Cd \mid\mid Hg_2^{2+} \mid Hg$
b) $Cd^{2+} \mid Hg_2^{2+} \mid\mid Cd \mid Hg$
c) $Cd \mid Cd^{2+} \mid\mid Hg_2^{2+} \mid Hg$
d) $Cd^{2+} \mid Hg \mid\mid Hg_2^{2+} \mid Cd$
e) $Hg \mid Cd \mid\mid Hg_2^{2+} \mid Cd^{2+}$

2. Consider an electrochemical cell where the following reaction takes place:

$$3Sn^{2+}(aq) + 2Al(s) \rightarrow 3Sn(s) + 2Al^{3+}(aq)$$

Which of the following is the correct cell notation for this cell?

a) $Al \mid Al^{3+} \mid\mid Sn^{2+} \mid Sn$
b) $Al^{3+} \mid Al \mid\mid Sn \mid Sn^{2+}$
c) $Sn \mid Sn^{2+} \mid\mid Al^{3+} \mid Al$
d) $Sn \mid Al^{3+} \mid\mid Al \mid Sn^{2+}$
e) $Al \mid Sn^{2+} \mid\mid Sn \mid Al^{3+}$

3. An early method of producing aluminum metal was the reaction of aluminum salts with sodium metal:

$Al^{3+} + 3Na(s) \rightleftharpoons Al(s)\ 3Na^{+}$ $\quad E° = +1.05$ V

What is $\Delta G°$ for this reaction

a) -304 kJ
b) -101 kJ
c) +101 kJ
d) +202 kJ
e) +304 kJ

4. Calculate ΔG for the following reaction:

$$I_2(s) + 2Br^{-}(aq) \rightarrow 2I^{-}(aq) + Br_2(l)$$

a) +105 kJ
b) -105 kJ e)
c) +312 kJ
d) +52 kJ
-312 kJ

5. If ΔG of the following reaction is -203 kJ, what is $E°$? $2Ag^{+}(aq) + Ni(s) \rightarrow 2Ag(s) + Ni^{2+}(aq)$

a) -1.05 V
b) +2.10 V
c) +0.0011 V
d) -0.011 V
e) +1.05 V

6. Given the two half reactions and their potentials, which net reaction is spontaneous?

$Mg^{2+}(aq) + 2e^{-} \rightarrow Mg(s)$ $\quad E° = -237$ V
$Ni^{2+}(aq) + 2e^{-} \rightarrow Ni(s)$ $\quad E° = -0.25$ V

a) $Ni(s) + Mg^{2+}(aq) \rightarrow Mg(s) + Ni^{2+}(aq)$
b) $Ni^{2+}(aq) + Mg(s) \rightarrow Mg^{2+}(aq) + Ni(s)$
c) $Ni(s) + Mg(s) \rightarrow Mg^{2+}(aq) + Ni^{2+}(aq)$
d) $Mg^{2+}(aq) + Ni^{2+}(aq) \rightarrow Mg(s) + Ni(s)$
e) $Mg^{2+}(aq) + Mg(s) \rightarrow Ni(s) + Ni^{2+}(aq)$

7. Calculate $E°$ for the following reaction:

$$Sn^{4+}(aq) + 2K(s) \rightarrow Sn^{2+}(aq) + 2K^{+}(aq)$$

a) +6.00 V
b) -3.08 V
c) +3.08 V
d) +2.78 V
e) -2.78 V

8. Calculate $E°$ for the following reaction:

$$2Al^{3+}(aq) + 3Cd(s) \rightarrow 2Al(s) + 3Cd^{2+}(aq)$$

a) -2.06 V
b) +4.52 V
c) +2.06 V
d) -4.52 V
e) -1.26 V

9. Using data from the reduction potential table and the reaction

$2Ag(s) + Pt^{2+}(aq) \rightarrow Pt(s) + 2Ag^{+}(aq)$ $\quad E° = 0.38$ V

calculate the standard reduction potential of the half-reaction

$$Pt^{2+}(aq) + 2e^{-} \rightarrow Pt(s)$$

a) -1.18 V
b) -0.40 V
c) 0.40 V
d) 1.18 V
e) 2.00 V

10. An electrochemical cell of notation $Pd \mid Pd^{2+} \mid\mid Cu^{2+} \mid Cu$ has an $E° = -0.65$ V. If we know that the standard reduction potential of Cu^{2+}/Cu is $E° = 0.34$ V, what is the standard reduction potential for Pd^{2+}/Pd?

a) -0.99 V
b) -0.31 V
c) +0.31 V
d) 0.62 V
e) +0.99 V

11. What is the equilibrium constant for the following reaction at 298 K?

$2Ag^+(aq) + 2I^-(aq) \rightarrow I_2(s) + 2Ag(s)$ $E° = +0.265$ V

a) 2.99×10^4
b) 9.04×10^8
c) 7.73×10^3
d) 87.9
e) 1.60×10^7

12. What is the equilibrium constant for the following reaction at 20C?

$Fe(s) + Cu^{2+}(aq) \rightarrow Fe^{2+}(aq) + Cu(s)$ $E° = +0.78$ V

a) 2.3×10^{26}
b) 6.9×10^{26}
c) 1.4×10^{27}
d) 1.8×10^{28}
e) 1.2×10^{-21}

13. What is the cell potential for

$3Sn^{4+}(aq) + 2Al(s) \rightarrow 3Sn^{2+}(aq) + 2Al^{3+}(aq)$

$E° = 1.81$ V when $[Sn^{4+}] = 1.0$, $[Sn^{2+}] = 1.0 \times 10^{-2}$, and $[Al^{3+}] = 1.5 \times 10^{-3}$ at 298 K.

a) 1.70 V
b) 1.76 V
c) 1.81 V
d) 1.86 V
e) 1.93 V

14. If the potential cell is +1.32 V at Q = 0.0969 with n = 2, what is the standard potential of the cell?

a) +1.35 V
b) +1.48 V
c) +1.31 V
d) +1.34 V
e) +1.29 V

15. Predict the product at the anode when electric current is passed through a solution of KI.

a) $I_2(l)$
b) $K^+(aq)$
c) $H_2(g)$
d) $K(s)$
e) $O_2(g)$

16. If electric current is passed through aqueous LiBr, the product at the cathode would be _____ and the product at the anode would be _____.

a) $H_2O(l)$, $Li^+(aq)$
b) $Br_2(l)$, $Li(s)$
c) $Li(s)$, $Br_2(l)$
d) $Br_2(l)$, $H_2(g)$
e) $H_2(g)$, $Br_2(l)$

17. How long would it take to deposit 1.36 g of copper from an aqueous solution of copper(II) sulfate by passing a current of two amperes through the solution?

a) 2070 sec
b) 1.11×10^{-5} sec
c) 2570 sec
d) 736 sec
e) 1030 sec

18. If a current of 6.0 amps is passed through a solution of Ag^+ for 1.5 hours, how many grams of silver are produced?

a) 0.60 g
b) 36 g
c) 0.34 g
d) 3.0 g
e) 1.0 g

19. How many kilowatt hours of electrical energy are required to plate 2.00 grams of silver from an aqueous solution of silver nitrate on to a necklace using 3.00V? (1 joule = 1 volt-coulomb and 1 kwh = 3.60×10^6 J)

a) 0.00135 kwh
b) 0.000165 kwh
c) 32.4 kwh
d) 0.00149 kwh
e) 2.07 kwh

20. How is aluminum currently produced in industry?

a) by reduction of Al^{3+} in Al_2O_3 with Na(s)
b) electrochemical reduction of pure Al_2O_3 to give Al and O_2
c) electrolysis of AlF_3 to give Al and F_2
d) electrolysis of a mixture of Al_2O_3 and Na_3AlF_6 to give Al and O_2
e) by reduction of Al^{3+} in Al_2O_3 with CO(g)

21. How as aluminum originally made?

a) the Hall-Heroult process
b) Al_2O_3 mixed with cryolite is electrolyzed
c) electrolysis of molten Al_2O_3
d) mining and purifying directly
e) reducing $AlCl_3$ with sodium

22. Using data from the reduction potential table, predict which of the following is the best oxidizing agent.

a) F_2
b) Ag
c) Sn^{4+}
d) Ag^+
e) Al^{3+}

23. Under acidic conditions the bromate ion is reduced to the bromide ion. Write the balance half-reaction for this process.

a) $BrO_3^- + 6H^+ + 6e \rightarrow Br^- + 3H_2O$
b) $2BrO_3^- + 6H^+ \rightarrow Br_2^- + 6H_2O + 3e$
c) $Bro_3^- + 6H_2O + 10e \rightarrow Br_2^- + 12H^+ + 3\ O_2$
d) $2BrO_3^- + 6H_2O \rightarrow 2Br^- + 12H^+ + 6\ O_2 + 8e$
e) $2BrO_3^- + 6H^+ \rightarrow Br_2^- + 3H_2O + 3e$

24. Balance the following redox edquation which occurs in acidic solution.

$N_2H_4(g) + BrO_3^-(aq) \rightarrow Br^-(aq) + N_2(g)$

a) $3N_2H_4 + BrO_3^- \rightarrow 3N_2 + Br^- + 3H_2O + 6H^+$
b) $N_2H_4 + BrO_3^- + 2H^+ \rightarrow 2Br^- + N_2 + 3H_2O$
c) $3N_2H_4 + 2BrO_3^- + 12H^+ \rightarrow 3N_2 + 2Br^- + 6H_2O + 12H^+$
d) $N_2H_4 + 2BrO_3^- + 8H^+ \rightarrow 2Br^- + N_2 + 6H_2O$
e) $3N_2H_4 + 2BrO_3^- \rightarrow 3N_2 + 2Br^- + 6H_2O$

25. Which of the following reactions is NOT a redox reaction?

a) $2HgO(s) \rightarrow 2\ Hg(l) + O_2(g)$
b) $H_2(g) + Br_2(g) \rightarrow 2HBr(g)$
c) $2HCl(aq) + Zn(s) \rightarrow H_2(g) + ZnCl_2(aq)$
d) $H_2CO_3(aq) \rightarrow H_2O(l) + CO_2(g)$
e) $2KClO_3 \rightarrow 2KCl(s) + 3\ O_2(g)$

Standard Reduction Potentials at 25°C E° (volts)

$F_2(g) + 2e^- \rightarrow 2F^-(aq)$	+2.87
$Au^{3+} + 3e^- \rightarrow Au(s)$	+1.50
$Cl_2(g) + 2e^- \rightarrow 2Cl^-(aq)$	+1.36
$O_2(g) + 4H_3O^+(aq) + 4e^- \rightarrow 6H_2O(l)$	+1.23
$Br_2(l) + 2e^- \rightarrow 2Br^-(aq)$	+1.08
$Ag^+(aq) + e^- \rightarrow Ag(s)$	+0.80
$I_2(s) + 2e^- \rightarrow 2I^-(aq)$	+0.535
$Cu^{2+}(aq) + 2e^- \rightarrow Cu(s)$	+0.337
$Sn^{4+}(aq) + 2e^- \rightarrow Sn^{2+}(aq)$	+0.15
$Sn^{2+}(aq) + 2e^- \rightarrow Sn(s)$	-0.14
$Cd^{2+}(aq) + 2e^- \rightarrow Cd(s)$	-0.40
$Zn^{2+}(aq) + 2e^- \rightarrow Zn(s)$	-0.763
$2H_2O(l) + 2e^- \rightarrow H_2(g) + 2OH^-(aq)$	-0.828
$Al^{3+}(aq) + 3e^- \rightarrow Al(s)$	-1.66
$K^+(aq) + e^- \rightarrow K(s)$	-2.93
$Li^+(aq) + e^- \rightarrow Li(s)$	-3.045

Answers:

1.	C	11.	B	21.	E
2.	A	12.	B	22.	A
3.	A	13.	E	23.	A
4.	A	14.	E	24.	E
5.	E	15.	A	25.	D
6.	B	16.	E		
7.	C	17.	A		
8.	E	18.	B		
9.	D	19.	D		
10.	E	20.	D		