South Pasadena • Chemistry

Name_____
Period ___ Date ___/___/___

19 • Acids and Bases

N E U T R A L I Z A T I O N

Write the dissociation equations for the following acids and bases:

1. $HCl \rightarrow$

2. $NaOH \rightarrow$

3. $HC_2H_3O_2 \rightarrow$

4. $Ba(OH)_2 \rightarrow$

5. $NH_3 + H_2O \rightarrow$

6. $H_2SO_4 \rightarrow$

Write balanced equations for the following neutralizations:

7. $HCl + NaOH \rightarrow$

8. $HC_2H_3O_2 + NaOH \rightarrow$

9. $HBr + Ba(OH)_2 \rightarrow$

10. $H_2SO_4 + NaOH \rightarrow$

11. $H_2SO_4 + Mg(OH)_2 \rightarrow$

12. $HCl + NH_3 \rightarrow$

Circle the acids and bases that are STRONG:

HCl	$NaOH$	$HC_2H_3O_2$	$Ba(OH)_2$
H_2SO_4	NH_3	HBr	$Mg(OH)_2$

The six strong acids are:

The eight strong bases are: